Test Booklet Code

English	50

Rank Pro

Re – NEET (UG)- 2026

|| DATE: 21-06-2026 ||

Important Instruction

1. The test is of **3 hours 15 minutes** duration and the Test Booklet contains **180** multiple-choice questions (four options with a single correct answer) from **Physics**, **Chemistry**, and **Biology** (**Botany and Zoology**). All questions are compulsory.

2. Each question carries **4 marks**. For **each correct response**, the candidate will get **4 marks**. **For each incorrect response**, **1 mark will be deducted** from the total scores. The maximum mark is **720**.

3. Use a **Blue/Black** ballpoint Pen only for writing particulars on this page/marking responses on the Answer Sheet.

4. Rough work is to be done in the space provided for this purpose in the Test Booklet only.

5. On completion of the test, the candidate **must hand over the Answer Sheet (ORIGINAL and OFFICE Copy)** to the Invigilator before leaving the room/hall. The candidates are allowed to take away this Test Booklet with them.

6. The CODE for this Booklet is **50**.

7. The candidates should ensure that the Answer Sheet is not folded. Do not make any stray marks on the Answer Sheet. Do not write your Roll No. anywhere else except in the specified space in the Test Booklet/Answer Sheet. The use of white fluid for correction is **NOT** permissible on the Answer Sheet.

8. Each candidate must show on-demand his/her Admit Card to the Invigilator.

9. No candidate, without special permission of the Centre Superintendent or Invigilator, would leave his/her seat.

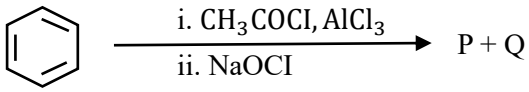
10. Use of an Electronic/Manual Calculator is prohibited.

11. The candidates are governed by all Rules and Regulations of the examination with regard to their conduct in the Examination Room/Hall. All cases of unfair means will be dealt with as per Rules and Regulations of this examination.

12. No part of the Test Booklet and Answer Sheet shall be detached under any circumstances.

13. The candidates will write the Correct Test Booklet Code as given in the Test Booklet/Answer Sheet in the Attendance Sheet.

46. For the following reaction sequence, choose the correct option.



(1) If P is the sodium salt of a carboxylic acid, Q is a primary alcohol.

(2) P and Q are aromatic compounds.

(3) If P gives a carboxylic acid on acidification, Q gives a poisonous gas on exposure to air and light.

(4) Both P and Q are carbonyl compounds.

47. Given below are two statements:

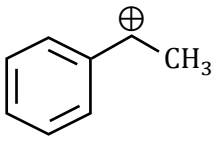
Statement-I: $[Fe(ox)_3]^{3-}$ is chiral.

Statement-II: trans – $[Cr(H_2O)_2(ox)_2]^-$ is chiral.
(Given: $oxH_2 = HOOC - COOH$)

In light of the above statements, choose the most appropriate answer from the options given below:

(1) Both Statement-I and Statement-II are correct
(2) Both Statement-I and Statement-II are incorrect
(3) Statement-I is correct but Statement-II is incorrect
(4) Statement-I is incorrect but Statement-II is correct

48. The following carbocation is stabilized by the interaction of the empty p orbital with



(1) filled σ and filled π orbitals

(2) empty σ and empty π^* orbitals

(3) empty σ^* and filled π orbitals

(4) empty σ^* and empty π^* orbitals

49. In potash alum, the ratio of K^+ and SO_4^{2-} ions is

(1) 1 : 2 (2) 2 : 1
(3) 2 : 3 (4) 3 : 2

50. The correct statement about peptides and proteins is

(1) Tertiary structure of proteins has two or more polypeptide subunits.

(2) Only the proteins having a quaternary structure are biologically active.

(3) In β-pleated sheet structures, peptide chains are held together by intermolecular hydrogen bonds.

(4) In α-helices, the polypeptide chain is twisted into a left-handed screw (helix) through intramolecular hydrogen bonds.

51. The numbers 17.0145 and 21.0235 were rounded to three figures after the decimal point. The resulting numbers, respectively, are

(1) 17.014 and 21.023
(2) 17.015 and 21.023
(3) 17.014 and 21.024
(4) 17.015 and 21.024

52. The correct order of solubility given salts in water at 298 K is

Salt	K_{sp} at 298 K
AgBr	5.0×10^{-13}
$Zn(OH)_2$	1.0×10^{-15}
Hg_2Cl_2	1.3×10^{-18}

(1) $Hg_2Cl_2 > Zn(OH)_2 > AgBr$
(2) $AgBr > Zn(OH)_2 > Hg_2Cl_2$
(3) $Hg_2Cl_2 > AgBr > Zn(OH)_2$
(4) $Zn(OH)_2 > AgBr > Hg_2Cl_2$

53. Among the following options, the correct trend in the electron gain enthalpy is

(1) F > Cl > Br > I

(2) Br > Cl > F > I

(3) Cl > F > Br > I

(4) I > Br > Cl > F

54. **Assertion (A):** For an ideal solution formed by mixing liquids P and Q, Δ_{mix} H = 0 and Δ_{mix} V = 0

Reason (R): No interactions occur between P and Q In the light of the above statements, choose the most appropriate answer from the options given below.

(1) Both A and R are correct and R is the correct explanation of A.

(2) Both A and R are correct but R is NOT the correct explanation of A.

(3) A is correct but R is not correct.

(4) A is not correct but R is correct.

55. The amino acid that gives a red-blood colour on treating its sodium fusion extract with sodium nitroprusside is

(1) leucine (2) threonine

(3) methionine (4) serine

56. The standard electrode potential (E°) for the half-cell reaction $Fe^{3+} + e^{-} \rightarrow Fe^{2+}$ at 298 K is

(Given: $E^{\circ}(Fe^{3+}/Fe) = -0.04\,V$ and $E^{\circ}(Fe^{2+}/Fe) = -0.44\,V$ at 298 K)

(1) +0.40 V (2) +0.76 V

(3) –0.48 V (4) +0.92 V

57. In an acidic medium, 10 mL of 0.25 M oxalic acid is titrated with $KMnO_4$ solution. If the volume of $KMnO_4$ solution required to reach end point is 10 mL, the strength of the $KMnO_4$ solution is

(1) 0.10 M (2) 0.20 M

(3) 0.25 M (4) 0.15 M

58. According to crystal field theory, the correct order of ligands with respect to their decreasing order of field strength is

(1) $CO > NH_3 > H_2O > Cl^-$

(2) $CO > H_2O > NH_3 > Cl^-$

(3) $Cl^- > H_2O > NH_3 > CO$

(4) $Cl^- > NH_3 > H_2O > CO$

59. Two moles of an ideal gas undergo free expansion from 10 L to 100 L at 300 K. The values of ΔS_{system} and $\Delta S_{surroundings}$ are (R is universal gas constant)

(1) $\Delta S_{system} = 0$; $\Delta S_{surroundings} = 0$

(2) $\Delta S_{system} = 4.606R$; $\Delta S_{surroundings} = -4.606R$

(3) $\Delta S_{system} = 0$; $\Delta S_{surroundings} = 4.606R$

(4) $\Delta S_{surroundings} = 4.606R$; $\Delta S_{system} = 0$

60. $2A \xrightarrow{k} B$ is a zero-order reaction, where k = 1.0 mol L^{-1} min^{-1}. If the initial concentration of A is 2 M, then the time taken to complete 75% of the reaction will be

(1) 1.5 min (2) 0.75 min

(3)1.0 min (4) 2.0 min

61. Given below are two statements: One is labelled as Assertion A and the other is labelled as Reason R.

Assertion A: Generally, 3d transition metals have high melting points.

Reason R: Involvement of 3d-electrons in addition to 4s-electrons in the interatomic metallic bonding.

In light of the above statements, choose the most appropriate answer from the options given below:

(1) Both A and R are correct and R is the correct explanation of A

(2) Both A and R are correct and R is NOT the correct explanation of A

(3) A is correct but R is not correct.

(4) A is not correct but R is correct

62. For a salt XY, which is a strong electrolyte, the plot of Λ_m versus $\sqrt{c}$ has a slope of -90.0 S cm^2 mol$^{-3/2}$ L$^{1/2}$ at 298 K. At 0.01 M concentration of XY, the value of Λ_m is 145.0 S cm^2 mol^{-1}. The limiting molar conductivity of Y^- ion ($\lambda^0_{Y^-}$, *in* S cm^2 mol^{-1}) at 298 K will be

(Given: $\lambda^0_{X^+} = 74.0\ S$ cm^2 mol^{-1})

(1) 80.0 (2) 100.0

(3) 90.0 (4) 76.0

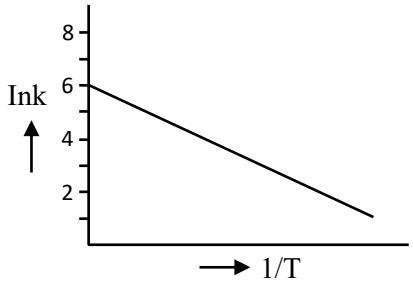
63. The amount of carbon dioxide evolved upon complete combustion of 116 g of n-butane is

(Given: atomic mass in amu H = 1, C = 12 and O = 16)

(1) 352 g (2) 322 g

(3) 176 g (4) 362 g

64. For an elementary chemical reaction, the Arrhenius plot is given below.



If the energy of activation is 6.64 kJ mol^{-1} and R = 8.3 J K^{-1} mol^{-1}, the temperature at which the rate constant becomes e^2 min^{-1}, is

(1) 125 K (2) 150 K

(3) 200 K (4) 250 K

65. Given below are two statements:

Statement-I: Heating NaCl with concentrated H_2SO_4 and MnO_2 results in oxidation of Mn.

Statement-II: Heating Nal with concentrated H_2SO_4 and MnO_2 results in reduction of Mn.

In light of the above statements, choose the most appropriate answer from the options given below.

(1) Both Statement-I and Statement-II are correct.

(2) Both Statement-I and Statement-II are incorrect.

(3) Statement-I is correct but Statement-II is incorrect.

(4) Statement-I is incorrect but Statement-II is correct.

66. Among the species given below, the spin-only magnetic moment is highest for

(Given: Atomic number of Ti = 22, Mn = 25, Fe = 26 and Co = 27)

(1) $[Mn(CN)_6]^{3-}$

(2) $[Fe(CN)_6]^{3-}$

(3) $[Co(NH_3)_6]^{3+}$

(4) $[Ti(H_2O)_6]^{3+}$

67. The lanthanide ion having four unpaired electrons is

(Given: Atomic numbers of Ce = 58, Nd = 60, Tb = 65 and Ho = 67)

(1) Nd^{3+} (2) Ce^{3+}

(3) Tb^{3+} (4) Ho^{3+}

68. The formula of tetraammineaquachloridocobalt(III) chloride is

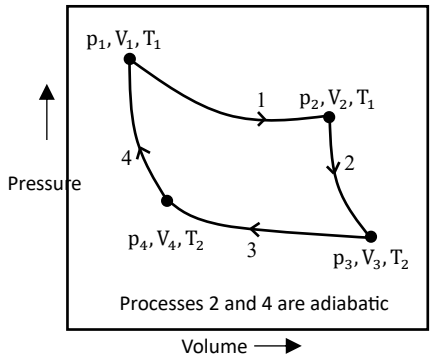
(1) $[Co(NH_3)_4Cl_2] \times H_2O$

(2) $[Co(NH_3)_4]Cl_3 \times H_2O$

(3) $[Co(NH_3)_4(H_2O)Cl]Cl$

(4) $[Co(NH_3)_4(H_2O)Cl]Cl_2$

69. Consider the reversible processes for 1.0 mol of an ideal gas as shown in the figure.



w_1, w_2, w_3 and w_4 represent work done (in calories) in the processes 1, 2, 3 and 4, respectively; ΔU_2 and ΔU_4 are changes in the internal energy for the processes 2 and 4, respectively.

[use R = 2 cal K^{-1} mol^{-1}]
The correct option is

(1) $w_1 + w_3 = -2T_1 \ln\frac{V_2}{V_1} - 2T_2 \ln\frac{V_4}{V_3}$

(2) $w_2 + w_4 = \Delta U_2 - \Delta U_2$

(3) $w_1 + w_2 = 2T_1 \ln\frac{V_2}{V_1}$

(4) $w_1 + w_2 + w_3 + w_4 = 0$

70. Given below are two statements: One is labelled as Assertion A and the other is labelled as Reason R.

Assertion A: The first ionization enthalpy of O is lower than that of N and F.

Reason R: The loss of an electron from O leads to stable half-filled *p* orbital.

In light of the above statements, choose the most appropriate answer from the options given below:

(1) Both A and R are correct and R is the correct explanation of A.

(2) Both A and R are correct and R is NOT the correct explanation of A.

(3) A is correct but R is not correct.

(4) A is not correct but R is correct.

71. Consider the following statements about the solutions formed by mixing two liquids.

A. An ideal solution thus formed obeys Raoult's law throughout the composition range.
B. Mixture of chloroform and acetone shows negative deviation from Raoult's law.
C. Mixture of aniline and phenol shows positive deviation from Raoult's law.

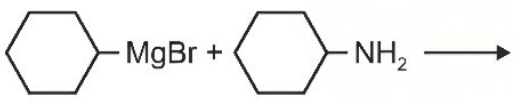
(1) A and B only

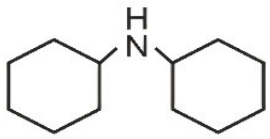
(2) B and C only

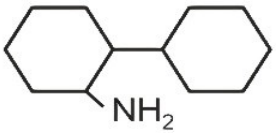
(3) A only

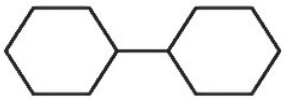
(4) A and C only

72. One of the products formed in the following reaction is



(1) 

(2) 

(3) 

(4) 

73. The correct statement is

(1) Boron has a maximum covalency of four.

(2) Beryllium has three valence orbitals.

(3) Magnesium has a maximum covalency of four.

(4) Aluminium has five valence orbitals.

74. A protein undergoes reversible thermal denaturation from its initial state N to denatured state D according to N ⇌ D. At 60°C, the concentrations of both N and D are equal at equilibrium, and the standard enthalpy change of denaturation is 666 kJ mol^{-1}. The standard entropy change (ΔS^{o} in kJ $K^{-1}mol^{-1}$) of the protein upon denaturation at 60°C is closest to

(1) 2.0

(2) 2000.0

(3) 333.0

(4) 11.1

75. Match the species in List-I with their geometry in List-II.

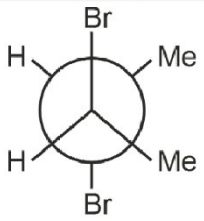
	List – I		List – II
A.	PCl_5	I.	Tetrahedral
B.	BrF_5	II.	Square Planar
C.	BF_4^-	III.	Trigonal bipyramidal
D.	$[Ni(CN)_4]^{2-}$	IV.	Square pyramidal

Choose the correct answer from the options:

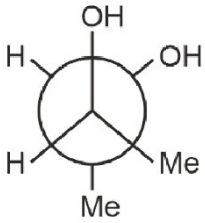
(1) A - IV, B - III, C - I, D - II
(2) A - III, B - IV, C - I, D - II
(3) A – III, B - I, C - II, D - IV
(4) A - III, B - II, C - I, D – IV

76. Given below are two statements:

Statement I: trans-But-2-ene upon treatment with Br_2 in CCl_4 gives the following product.



Statement II: cis-But-2-ene upon treatment with alkaline $KMnO_4$ gives the following product.



In the light of the above statements, choose the most appropriate answer from the options given below.

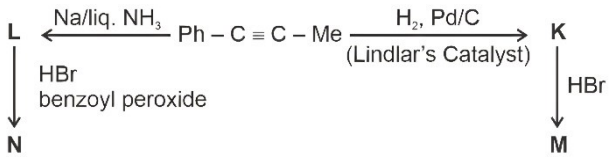
(1) Both Statement I and Statement II are correct

(2) Both Statement I and Statement II are incorrect

(3) Statement I is correct but Statement II is incorrect

(4) Statement I is incorrect but Statement II is correct

77. Consider the following reaction sequences and choose the correct option.



(1) K and L are geometrical isomers

(2) K and L are enantiomers

(3) M and N are geometrical isomers

(4) M and N are stereoisomers

78. The complex which has facial and meridional isomers is
(Given: py = pyridine and en = $H_2N - CH_2 - CH_2 - NH_2$)

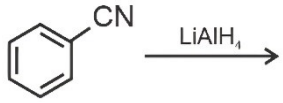
(1) $[Cr(py)_3(Cl)_3]$

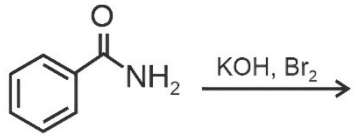
(2) $[Cr(H_2O)_6]^{3+}$

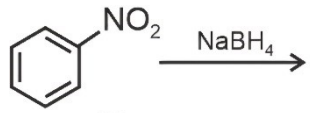
(3) $[Co(NH_3)_4(H_2O)_2]^{3+}$

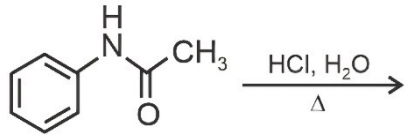
(4) $[Ni(en)_2(H_2O)_2]^{2+}$

79. Identify the reactions which give aniline as the major product.

A. 

B. 

C. 

D. 

Choose the correct answer from the options given below.

(1) A and B only

(2) B and D only

(3) A and C only

(4) C and D only

80. Match the vitamins in List I with their sources in List II

	List I		List II
A.	vitamin A	I.	meat
B.	vitamin B_{12}	II.	sunflower oil
C.	vitamin E	III.	green leafy vegetables
D.	vitamin K	IV.	carrots

Choose the correct answer from the options given below.

(1) A-II, B-III, C-IV, D-I

(2) A-IV, B-I, C-II, D-III

(3) A-IV, B-II, C-I, D-III

(4) A-III, B-I, C-IV, D-II

81. The correct decreasing order of oxidation state of the underlined atom in each molecule is

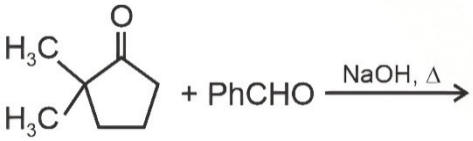
(1) $\underline{P_4O_{10}} > \underline{S}O_3 > H_2\underline{O}$

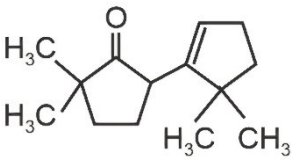
(2) $\underline{N_2}O_5 > \underline{Al_2}O_5 > H_2\underline{S}$

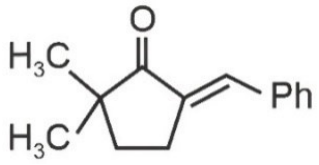
(3) $\underline{Pb}O_2 > \underline{N_2}O_3 > \underline{S}O_3$

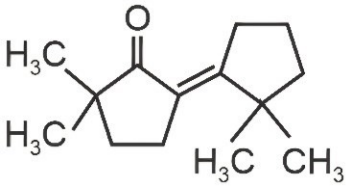
(4) $\underline{P_4}O_6 > \underline{Cl_2}O_7 > \underline{Al}H_3$

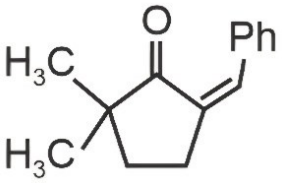
82. The compound that CANNOT be obtained from the aldol condensation reaction shown below, is



(1) 

(2) 

(3) 

(4) 

83. Among the following, the compound having conjugated double bonds is

(1) hepta-1,3-diene
(2) hepta-1,4-diene
(3) hepta-1,5-diene
(4) hepta-1,6-diene

84. Given below are two statements:

Statement-I: Oxidation of p-nitrotoluene with acidic $KMnO_4$ gives an acid that is stronger than benzoic acid.

Statement-II: Reduction of p-nitrotoluene with Sn/HCl followed by neutralization gives an amine that is more basic than aniline.

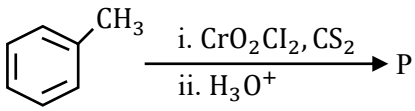
In light of the above statements, choose the most appropriate answer from the options given below.

(1) Both Statement-I and Statement-II are correct
(2) Both Statement-I and Statement-II are incorrect
(3) Statement-I is correct but Statement-II is incorrect
(4) Statement-I is incorrect but Statement-II is correct

85. The green paramagnetic species formed by heating $KMnO_4$ at 513 K is

(1) K_2MnO_4
(2) Mn_3O_4
(3) MnO
(4) KO_2

86. Consider the following reaction, and choose the correct option.



(1) On treating compound P with saturated $NaHCO_3$ solution, brisk effervescence is observed

(2) Compound P can be prepared by treating benzene with anhydrous $AlCl_3$ and CH_3COCl

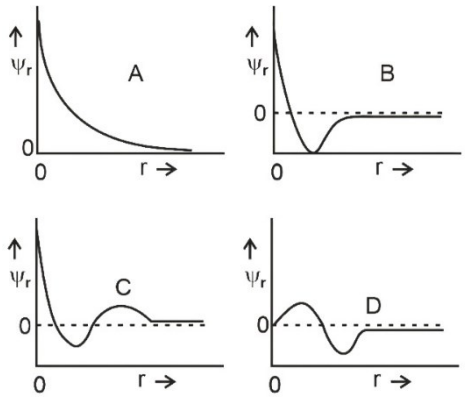
(3) On treatment with bromine water, compound P gives a white precipitate

(4) Compound P is obtained by the hydrogenation of benzoyl chloride with Pd on $BaSO_4$

87. A 1:3 electrolyte in an aqueous solution is

(1) $[CoCl_2(NH_3)_4]Cl$
(2) $[CoCl(NH_3)_5]Cl_2$
(3) $[Co(NH_3)_6]Cl_3$
(4) $[Co(NH_3)_3(NO_2)_3]$

88. Consider the following schematic plots of orbital wavefunction (ψr) against distance (r) from the nucleus.



The figure representing two radial nodes in the orbital is

(1) A (2) B

(3) C (4) D

89. Arrange the following compounds in the increasing order of polarity

A. $CH_3CH_2OCH_2CH_3$

B. CH_3CH_2OH

C. CH_3COCH_3

D. CH_3COOH

Choose the correct answer from the options given below.

(1) A < B < C < D (2) C < A < D < B

(3) C < A < B < D (4) A < C < B < D

90. The highest occupied molecular orbital for Ne_2 is

(1) π_{2_p} (2) σ_{2_p}

(3) $\pi^*_{2_p}$ (4) $\sigma^*_{2_p}$